

Product Information

4-Nitroaniline

Product Number **N 2128**
Store at Room Temperature

Product Description

Molecular Formula: C₆H₆N₂O₂

Molecular Weight: 138.1

CAS Number: 100-01-6

Melting Point: 146 °C¹

Synonyms: p-nitroaniline, p-nitroaniline¹

p-Nitroaniline is a chromogenic molecule that is used as a dyestuff intermediate in industrial applications.¹ In biochemical research, enzyme assays utilize modified aminoacyl or peptidyl p-nitroanilines as substrates. The enzyme catalyzes the release of free p-nitroaniline, which is the basis of the colorimetric determination of the enzyme activity.^{2,3,4}

Measurements are commonly performed at 410 nm ($E^{mM} = 8.80$), because of the absorbance overlap of the substrate and product at lower wavelengths.

Applications of the colorimetric properties of p-nitroanilide include the design of biopolymer drug delivery systems and of solid supports for enzyme immobilization.^{5,6,7} A kinetic analysis of the α -chymotrypsin catalyzed hydrolysis of aminoacyl and peptidyl p-nitroanilide substrates in vesicles has been reported.⁸

Computational studies on the crystal charge density of p-nitroaniline have been published.^{9,10}

Precautions and Disclaimer

For Laboratory Use Only. Not for drug, household or other uses.

Preparation Instructions

This product is soluble in ethanol (50 mg/ml), with heat as needed, yielding a slightly hazy, yellow to orange solution. It is soluble in mineral acids such as HCl, and also in alcohol (40 mg/ml) and ether (33 mg/ml).¹

References

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6. Shen, E., et al., Microphase separation in bioerodible copolymers for drug delivery. *Biomaterials*, **22**(3), 201-210 (2001).
7. Percot, A., et al., Immobilization of lipid vesicles on polymer support via an amphiphilic peptidic anchor: application to a membrane enzyme. *Bioconjug. Chem.*, **11**(5), 674-678 (2000).
8. Blocher, M., et al., Modeling of enzymatic reactions in vesicles: the case of α -chymotrypsin. *Biotechnol. Bioeng.*, **62**(1), 36-43 (1999).
9. Volkov, A., et al., Evaluation of net atomic charges and atomic and molecular electrostatic moments through topological analysis of the experimental charge density. *Acta Crystallogr. A*, **56**(Pt 3), 252-258 (2000).
10. Volkov, A., et al., On the origin of topological differences between experimental and theoretical crystal charge densities. *Acta Crystallogr. A*, **56**(Pt 4), 332-339 (2000).

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